

UNIVERSITY OF MYSORE
Postgraduate Entrance Examination September - 2023



**QUESTION PAPER
BOOKLET NO.**

Entrance Reg. No.

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SUBJECT CODE : 26

QUESTION BOOKLET

(Read carefully the instructions given in the Question Booklet)

COURSE : M.Sc.

**SUBJECT : GROUP-4 : CHEMISTRY/ORGANIC
CHEMISTRY**

MAXIMUM MARKS : 50

MAXIMUM TIME : 75 MINUTES

(Including time for filling O.M.R. Answer sheet)

INSTRUCTIONS TO THE CANDIDATES

1. The sealed question paper booklet containing 50 questions enclosed with O.M.R. Answer Sheet is given to you.
2. Verify whether the given question booklet is of the same subject which you have opted for examination.
3. Open the question paper seal carefully and take out the enclosed O.M.R. Answer Sheet outside the question booklet and fill up the general information in the O.M.R. Answer sheet. If you fail to fill up the details in the form as instructed, you will be personally responsible for consequences arising during evaluating your Answer Sheet.
4. During the examination:
 - a) Read each question carefully.
 - b) Determine the Most appropriate/correct answer from the four available choices given under each question.
 - c) Completely darken the relevant circle against the Question in the O.M.R. Answer Sheet. For example, in the question paper if "C" is correct answer for Question No.8, then darken against Sl. No.8 of O.M.R. Answer Sheet using Blue/Black Ball Point Pen as follows:

Question No. 8. (A) (B) ● (D) (Only example) (Use Ball Pen only)
5. Rough work should be done only on the blank space provided in the Question Booklet. Rough work should not be done on the O.M.R. Answer Sheet.
6. If more than one circle is darkened for a given question, such answer is treated as wrong and no mark will be given. See the example in the O.M.R. Sheet.
7. The candidate and the Room Supervisor should sign in the O.M.R. Sheet at the specified place.
8. Candidate should return the original O.M.R. Answer Sheet and the university copy to the Room Supervisor after the examination.
9. Candidate can carry the question booklet and the candidate copy of the O.M.R. Sheet.
10. The calculator, pager and mobile phone are not allowed inside the examination hall.
11. If a candidate is found committing malpractice, such a candidate shall not be considered for admission to the course and action against such candidate will be taken as per rules.
12. Candidates have to get qualified in the respective entrance examination by securing a minimum of 8 marks in case of SC/ST/Cat-I Candidates, 9 marks in case of OBC Candidates and 10 marks in case of other Candidates out of 50 marks.

INSTRUCTIONS TO FILL UP THE O.M.R. SHEET

1. There is only one most appropriate/correct answer for each question.
2. For each question, only one circle must be darkened with BLUE or BLACK ball point pen only. Do not try to alter it.
3. Circle should be darkened completely so that the alphabet inside it is not visible.
4. Do not make any unnecessary marks on O.M.R. Sheet.
5. Mention the number of questions answered in the appropriate space provided in the O.M.R. sheet otherwise O.M.R. sheet will not be subjected for evaluation.

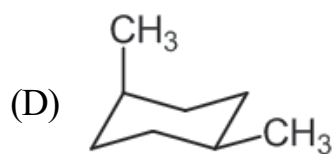
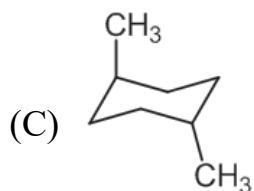
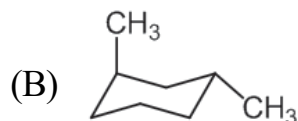
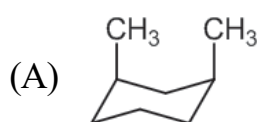
ಗಮನಿಸಿ : ಸೂಚನೆಗಳ ಕನ್ನಡ ಆವೃತ್ತಿಯು ಈ ಪುಸ್ತಕದ ಹಿಂಭಾಗದಲ್ಲಿ ಮುದ್ರಿಸಲ್ಪಟ್ಟಿದೆ.

- A plant cell shrinks when it is kept in
 - Isotonic solution with cell sap
 - Hypotonic solution
 - Water
 - Hypertonic solution
- If the activity of a solution is 0.82 and its concentration is 0.98 M, the activity coefficient is,
 - 0.988
 - 0.12
 - 0.9
 - 0.836
- The root mean square velocity (u), the average velocity (p) and most probable velocity (q) of a molecule are related to each other. Then which of the following is true?
 - $q > p > u$
 - $p > q > u$
 - $u > p > q$
 - $u = p = q$
- The frequency (cm^{-1}) of spectral lines in the rotational spectrum of a rigid diatomic molecule is obtained by
 - $BJ(J+1)$
 - $2B(J-1)$
 - $2BJ$
 - $2B(J+1)^2$
- Addition of Ag^+ , Pb^{2+} , Fe^{3+} and Si^{4+} causes coagulation of negatively charged colloidal sol. Then which of the following is true?
 - $\text{Fe}^{3+} > \text{Si}^{4+} > \text{Pb}^{2+} > \text{Ag}^+$
 - $\text{Ag}^+ < \text{Pb}^{2+} < \text{Fe}^{3+} < \text{Si}^{4+}$
 - $\text{Ag}^+ > \text{Pb}^{2+} > \text{Fe}^{3+} > \text{Si}^{4+}$
 - $\text{Ag}^+ = \text{Pb}^{2+} = \text{Fe}^{3+} = \text{Si}^{4+}$
- A container of volume 5.0 L is divided into two compartments of equal size. In the left compartment there is nitrogen at 1.0 atmosphere pressure and 25°C , in the right compartment there is hydrogen at the same temperature and pressure. What will happen when the partition is removed?
 - The entropy increases, and the free energy decreases
 - The entropy decreases, and the free energy decreases
 - The entropy increases, and the free energy increases
 - The entropy decreases, and the free energy increases

7. A unit cell has the following characteristics, $a \neq b \neq c$; $\alpha = \gamma = 90^\circ$, $\beta \neq 90^\circ$. The unit cell belongs to the crystal system
- (A) Orthorhombic (B) Rhombohedral
(C) Monoclinic (D) Triclinic
8. The triple point in the phase diagram of one component system has
- (A) One degree of freedom (B) Two degrees of freedom
(C) Three degrees of freedom (D) Zero degree of freedom
9. Which of the following statements associated with corrosion is incorrect?
- (A) Iron corrodes more readily than aluminium
(B) Cathodic protection prevents corrosion by using a sacrificial anode
(C) A cathodic metal has both anodic and cathodic areas
(D) Corrosion involves both oxidation and reduction
10. The low quantum efficiency of photochemical reactions are due to
- (A) Deactivation of the reacting molecules
(B) Dissolution of the reacting molecule
(C) Fragmentation of the products
(D) Formation of stable products
11. The solubility 's' of a sparingly soluble salt is related to its equivalent conductance at infinite dilution λ_∞ by the relation (k is specific conductance)
- (A) $s = \frac{k \times 1000}{\lambda_\infty - \lambda}$ (B) $s = \frac{c \times 1000}{\lambda_\infty - \lambda}$
(C) $s = \frac{k \times 1000}{\lambda_\infty}$ (D) $s = \frac{c \times 1000}{\lambda_\infty}$
12. The molecular weight of NaCl, assuming 100% dissociation in solution, as determined by elevation of boiling point method is
- (A) equal to 58.5 (B) 29.25
(C) greater than 58.5 (D) Zero

13. The conductivity of 0.01 M NaCl solution is $0.00147 \text{ ohm}^{-1} \text{ cm}^{-1}$. When 50 ml of water is added to the above solution, its conductivity
- (A) Increases
 (B) Decreases
 (C) Remains unchanged
 (D) First increases and then decreases
14. The rate constant for a first order reaction is 60 s^{-1} . The time taken to reduce the initial concentration of the reactant to its $1/16^{\text{th}}$ value will be
- (A) 0.00462 s
 (B) 0.462 s
 (C) 0.0462 s
 (D) 4.63 s
15. In a $\text{H}_2\text{-O}_2$ fuel cell, combustion of hydrogen occurs to
- (A) Remove adsorbed oxygen from electrode surface
 (B) Create potential difference between the two electrodes
 (C) Produce high purity water
 (D) Generate Heat
16. Which term describes the closeness of a measured value to the true value?
- (A) Accuracy
 (B) Precision
 (C) Error
 (D) Standard deviation

17. The least stable molecule is



18. Which chromatography technique is commonly used for the separation and quantification of volatile organic compounds?

- (A) Thin-Layer Chromatography
- (B) High-Performance Liquid Chromatography
- (C) Gas Chromatography
- (D) Ion Chromatography

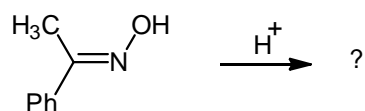
19. A solution of quinine sulphate on exposure to visible light exhibits

- (A) Fluorescence
- (B) Phosphorescence
- (C) Chemiluminescence
- (D) None of the above

20. Error mainly caused by human mistake is

- (A) Gross error
- (B) Random error
- (C) Absolute error
- (D) Instrumental error

21. The major product formed in the following reaction is



- (A) N-methyl benzamide
- (B) Acetanilide
- (C) Benzanilide
- (D) Benzoic acid

22. Heterocyclic rings present in nicotine are

- (A) Pyridine and pyrrole
- (B) Pyridine and pyrrolidine
- (C) Piperidine and pyrrole
- (D) Piperidine and pyrrolidine

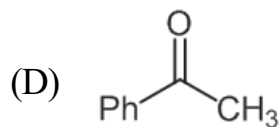
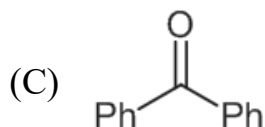
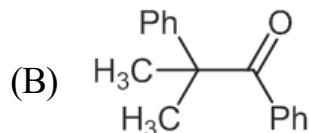
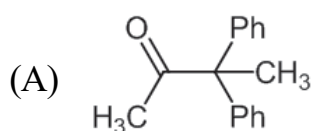
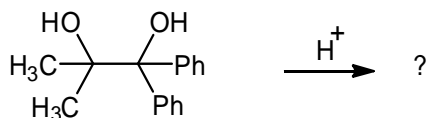
23. Octane-2,7-dione is obtained from the ozonolysis of

- (A) 2,3-dimethyl cyclopentene
- (B) 1,2-dimethyl cyclopentane
- (C) 1,2-dimethyl cyclopentene
- (D) 1,2-dimethyl cyclohexene

24. Which one of the following bio-molecule contains non-transition metal?

- (A) Vitamin B12 (B) Haemoglobin
(C) Chlorophyll (D) Insulin

25. The product formed in the following reaction is



26. Which one of the following amino acid is non essential?

- (A) Lysine (B) Thryptophan
(C) Histidine (D) Proline

27. Nylon is an example of

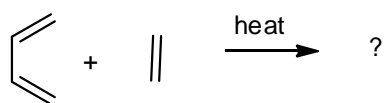
- (A) Polyester (B) Polyamide
(C) Polysaccharide (D) Polythene

28. Name the product(s) in the following reaction.



- (A) 1-Phenylethanol (B) Benzyl alcohol
(C) Diphenylmethanol (D) Benzoic acid

29. Name the product of the following reaction



- (A) Cyclohexane (B) Cyclopentane
(C) Cyclohexene (D) Cyclobutene

30. Which of the following molecule is more acidic?

- (A) CCl_3COOH (B) CH_3COOH
(C) ClCH_2COOH (D) Cl_2CHCOOH

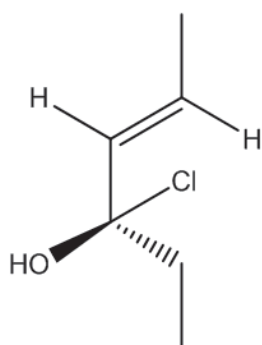
31. The least stable structural conformation of cyclohexane is

- (A) Chair form (B) Boat form
(C) Twist boat form (D) Half Chair form

32. Name the reaction which involves hydride ion transfer during the reaction

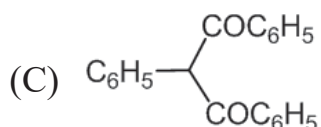
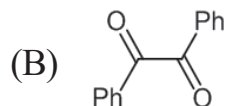
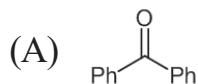
- (A) Cannizarro reaction (B) Aldol reaction
(C) Fries rearrangement (D) Classen reaction

33. IUPAC name for the following compound is



- (A) 3R,3(E)-Chlorohex-4-en-3-ol
(B) 4R,4(E)-Chlorohex-2-en-4-ol
(C) 3R,3(Z)-Chlorohex-4-en-3-ol
(D) 4S,4(E),Chlorohex-2-en-ol

34. The product in the following reaction Benzoyl chloride with benzene in the presence of aluminium chloride



(D) None of these

35. For quantum numbers $n=4$ and $l=3$ the maximum number of electrons in sub shell is

(A) 6

(B) 10

(C) 14

(D) 16

36. A group of four students A, B, C and D has arranged the covalent character of alkali metal halides in different order. One of the student has provided the correct order of covalent character is

(A) $\text{NaF} < \text{NaBr} < \text{NaI} < \text{NaCl}$

(B) $\text{NaF} < \text{NaCl} < \text{NaBr} < \text{NaI}$

(C) $\text{NaCl} < \text{NaBr} < \text{NaI} < \text{NaF}$

(D) $\text{NaI} < \text{NaBr} < \text{NaF} < \text{NaCl}$

37. Arvi, Atul, Ashok and Arjun had arranged in increasing order of the polarizing power of the following cationic species. Who had provided the correct order?

(A) Arvi : $\text{Ca}^{2+} < \text{Mg}^{2+} < \text{Be}^{2+} < \text{K}^+$

(B) Atul : $\text{Mg}^{2+} < \text{Be}^{2+} < \text{K}^+ < \text{Ca}^{2+}$

(C) Ashok : $\text{Be}^{2+} < \text{K}^+ < \text{Ca}^{2+} < \text{Mg}^{2+}$

(D) Arjun : $\text{K}^+ < \text{Ca}^{2+} < \text{Mg}^{2+} < \text{Be}^{2+}$

38. The bond between atoms of two elements with atomic numbers 11 and 17 is

(A) Metallic

(B) Electrocovalent

(C) Covalent

(D) Coordinate

39. One of the following compound acts as a reducing agent and an oxidizing agent is

(A) Na_2O

(B) NaNO_3

(C) KNO_3

(D) Na_2O_2

40. Liquid ammonia is used in refrigeration because
- (A) It has a high heat of vaporization (B) It has a high dipole moment
(C) It is a basic in nature (D) It is more volatile
41. The covalent nature of AlCl_3 can be justified on the basis of
- (A) Resonance (B) Fajan's rule
(C) Hund's rule (D) MO Theory
42. The color of the Cerium(IV) is due to
- (A) f-f transitions
(B) d-d transitions
(C) The presence of unpaired electrons
(D) The charge transfer
43. In a nuclear reactor, the function of moderator is
- (A) To slow down the speed of neutron
(B) To increase the speed of neutron
(C) To produce more neutrons
(D) To stop the nuclear reactor
44. Ankitha, Anvitha, Anusha and Anjali have performed the gravimetric analysis of nickel with dimethylglyoxime in a slight alkaline medium to give red colored complex that had provided consent composition the complex?
- (A) Ankitha: 1:1 Composition (B) Anvitha: 1:2 Composition
(C) Anusha: 1:3 Composition (D) Anjali: 1:4 Composition
45. The molecular formula of the complex tris(ethylene diammine) cobalt (III) sulphate is
- (A) $[\text{Co}(\text{en})_2\text{SO}_4]$ (B) $[\text{Co}(\text{en})_3\text{SO}_4]$
(C) $[\text{Co}(\text{en})_3]\text{SO}_4$ (D) $[\text{Co}(\text{en})_3]_2(\text{SO}_4)_3$

46. The geometries of the ammonia complexes of Ni^{2+} , Pt^{2+} and Zn^{2+} respectively are
- (A) Octahedral, square planar and tetrahedral
 - (B) Square planar, tetrahedral and octahedral
 - (C) Tetrahedral, octahedral and square planar
 - (D) Square planar, octahedral and tetrahedral
47. The Δ_0 value for $[\text{CoCl}_6]^{4-}$ is $18,000 \text{ cm}^{-1}$, then the value of Δt for $[\text{CoCl}_4]^{2-}$ would be
- (A) $18,000 \text{ cm}^{-1}$
 - (B) $12,000 \text{ cm}^{-1}$
 - (C) $8,000 \text{ cm}^{-1}$
 - (D) $4,000 \text{ cm}^{-1}$
48. Which one of the following exhibits a different property from rest of the group?
- (A) $[\text{Zn}(\text{NH}_3)_6]^{2+}$
 - (B) $[\text{Ni}(\text{NH}_3)_6]^{3+}$
 - (C) $[\text{Cr}(\text{NH}_3)_6]^{3+}$
 - (D) $[\text{Co}(\text{NH}_3)_6]^{3+}$
49. The non-sulfide ore is
- (A) Galena
 - (B) Malachite
 - (C) Cinnabar
 - (D) Copper pyrites
50. In an analysis of hematite ore for the determination of iron using cerium solution as titrant, the indicator used in this titrimetric analysis is
- (A) Diphenyl ammine
 - (B) Ferroin
 - (C) Eriochrome Black-T
 - (D) Patton and Reeders



Rough Work

ಅಭ್ಯರ್ಥಿಗಳಿಗೆ ಸೂಚನೆಗಳು

1. ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯ ಜೊತೆಗೆ 50 ಪ್ರಶ್ನೆಗಳನ್ನು ಹೊಂದಿರುವ ಮೊಹರು ಮಾಡಿದ ಪ್ರಶ್ನೆ ಪುಸ್ತಕವನ್ನು ನಿಮಗೆ ನೀಡಲಾಗಿದೆ.
2. ಕೊಟ್ಟಿರುವ ಪ್ರಶ್ನೆ ಪುಸ್ತಕವು, ನೀವು ಪರೀಕ್ಷೆಗೆ ಆಯ್ಕೆ ಮಾಡಿಕೊಂಡಿರುವ ವಿಷಯಕ್ಕೆ ಸಂಬಂಧಿಸಿದ್ದೇ ಎಂಬುದನ್ನು ಪರಿಶೀಲಿಸಿರಿ.
3. ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯ ಮೊಹರು ಜಾಗ್ರತೆಯಿಂದ ತೆರೆಯಿರಿ ಮತ್ತು ಪ್ರಶ್ನೆಪತ್ರಿಕೆಯಿಂದ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯನ್ನು ಹೊರಗೆ ತೆಗೆದು, ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯಲ್ಲಿ ಸಾಮಾನ್ಯ ಮಾಹಿತಿಯನ್ನು ತುಂಬಿರಿ. ಕೊಟ್ಟಿರುವ ಸೂಚನೆಯಂತೆ ನೀವು ನಮೂನೆಯಲ್ಲಿನ ವಿವರಗಳನ್ನು ತುಂಬಲು ವಿಫಲರಾದರೆ, ನಿಮ್ಮ ಉತ್ತರ ಹಾಳೆಯ ಮೌಲ್ಯಮಾಪನ ಸಮಯದಲ್ಲಿ ಉಂಟಾಗುವ ಪರಿಣಾಮಗಳಿಗೆ ವೈಯಕ್ತಿಕವಾಗಿ ನೀವೇ ಜವಾಬ್ದಾರಾಗಿರುತ್ತೀರಿ.
4. ಪರೀಕ್ಷೆಯ ಸಮಯದಲ್ಲಿ:
 - a) ಪ್ರತಿಯೊಂದು ಪ್ರಶ್ನೆಯನ್ನು ಜಾಗ್ರತೆಯಿಂದ ಓದಿರಿ.
 - b) ಪ್ರತಿ ಪ್ರಶ್ನೆಯ ಕೆಳಗೆ ನೀಡಿರುವ ನಾಲ್ಕು ಲಭ್ಯ ಆಯ್ಕೆಗಳಲ್ಲಿ ಅತ್ಯಂತ ಸರಿಯಾದ/ ಸೂಕ್ತವಾದ ಉತ್ತರವನ್ನು ನಿರ್ಧರಿಸಿ.
 - c) ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿನ ಸಂಬಂಧಿಸಿದ ಪ್ರಶ್ನೆಯ ವೃತ್ತಾಕಾರವನ್ನು ಸಂಪೂರ್ಣವಾಗಿ ತುಂಬಿರಿ. ಉದಾಹರಣೆಗೆ, ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯಲ್ಲಿ ಪ್ರಶ್ನೆ ಸಂಖ್ಯೆ 8ಕ್ಕೆ "C" ಸರಿಯಾದ ಉತ್ತರವಾಗಿದ್ದರೆ, ನೀಲಿ/ಕಪ್ಪು ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್ ಬಳಸಿ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯ ಕ್ರಮ ಸಂಖ್ಯೆ 8ರ ಮುಂದೆ ಈ ಕೆಳಗಿನಂತೆ ತುಂಬಿರಿ:
 ಪ್ರಶ್ನೆ ಸಂಖ್ಯೆ 8. (A) (B) (C) (D) (ಉದಾಹರಣೆ ಮಾತ್ರ) (ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್ ಮಾತ್ರ ಉಪಯೋಗಿಸಿ)
5. ಉತ್ತರದ ಪೂರ್ವಸಿದ್ಧತೆಯ ಬರವಣಿಗೆಯನ್ನು (ಚಿತ್ತು ಕೆಲಸ) ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯಲ್ಲಿ ಒದಗಿಸಿದ ಖಾಲಿ ಜಾಗದಲ್ಲಿ ಮಾತ್ರವೇ ಮಾಡಬೇಕು (ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯಲ್ಲಿ ಮಾಡಬಾರದು).
6. ಒಂದು ನಿರ್ದಿಷ್ಟ ಪ್ರಶ್ನೆಗೆ ಒಂದಕ್ಕಿಂತ ಹೆಚ್ಚು ವೃತ್ತಾಕಾರವನ್ನು ಗುರುತಿಸಲಾಗಿದ್ದರೆ, ಅಂತಹ ಉತ್ತರವನ್ನು ತಪ್ಪು ಎಂದು ಪರಿಗಣಿಸಲಾಗುತ್ತದೆ ಮತ್ತು ಯಾವುದೇ ಅಂಕವನ್ನು ನೀಡಲಾಗುವುದಿಲ್ಲ. ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿನ ಉದಾಹರಣೆ ನೋಡಿ.
7. ಅಭ್ಯರ್ಥಿ ಮತ್ತು ಕೊಠಡಿ ಮೇಲ್ವಿಚಾರಕರು ನಿರ್ದಿಷ್ಟಪಡಿಸಿದ ಸ್ಥಳದಲ್ಲಿ ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯ ಮೇಲೆ ಸಹಿ ಮಾಡಬೇಕು.
8. ಅಭ್ಯರ್ಥಿಯು ಪರೀಕ್ಷೆಯ ನಂತರ ಕೊಠಡಿ ಮೇಲ್ವಿಚಾರಕರಿಗೆ ಮೂಲ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆ ಮತ್ತು ವಿಶ್ವವಿದ್ಯಾನಿಲಯದ ಪ್ರತಿಯನ್ನು ಹಿಂದಿರುಗಿಸಬೇಕು.
9. ಅಭ್ಯರ್ಥಿಯು ಪ್ರಶ್ನೆ ಪುಸ್ತಕವನ್ನು ಮತ್ತು ಓ.ಎಂ.ಆರ್. ಅಭ್ಯರ್ಥಿಯ ಪ್ರತಿಯನ್ನು ತಮ್ಮ ಜೊತೆ ತೆಗೆದುಕೊಂಡು ಹೋಗಬಹುದು.
10. ಕ್ಯಾಲ್ಕುಲೇಟರ್, ಪೇಜರ್ ಮತ್ತು ಮೊಬೈಲ್ ಫೋನ್‌ಗಳನ್ನು ಪರೀಕ್ಷಾ ಕೊಠಡಿಯ ಒಳಗೆ ಅನುಮತಿಸಲಾಗುವುದಿಲ್ಲ.
11. ಅಭ್ಯರ್ಥಿಯು ದುಷ್ಕೃತ್ಯದಲ್ಲಿ ತೊಡಗಿರುವುದು ಕಂಡುಬಂದರೆ, ಅಂತಹ ಅಭ್ಯರ್ಥಿಯನ್ನು ಕೋರ್ಸ್‌ಗೆ ಪರಿಗಣಿಸಲಾಗುವುದಿಲ್ಲ ಮತ್ತು ನಿಯಮಗಳ ಪ್ರಕಾರ ಅಂತಹ ಅಭ್ಯರ್ಥಿಯ ವಿರುದ್ಧ ಕ್ರಮ ಕೈಗೊಳ್ಳಲಾಗುವುದು.
12. ಈ ಪ್ರವೇಶ ಪರೀಕ್ಷೆಯಲ್ಲಿ ಅರ್ಹರಾಗಲು ಒಟ್ಟು 50 ಅಂಕಗಳಲ್ಲಿ SC/ST/Cat-I ಅಭ್ಯರ್ಥಿಗಳು ಕನಿಷ್ಠ 8 ಅಂಕಗಳನ್ನು, OBC ಅಭ್ಯರ್ಥಿಗಳು ಕನಿಷ್ಠ 9 ಅಂಕಗಳನ್ನು ಮತ್ತು ಇನ್ನಿತರ ಅಭ್ಯರ್ಥಿಗಳು ಕನಿಷ್ಠ 10 ಅಂಕಗಳನ್ನು ಪಡೆಯತಕ್ಕದ್ದು.

ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯನ್ನು ತುಂಬಲು ಸೂಚನೆಗಳು

1. ಪ್ರತಿಯೊಂದು ಪ್ರಶ್ನೆಗೆ ಒಂದೇ ಒಂದು ಅತ್ಯಂತ ಸೂಕ್ತವಾದ/ಸರಿಯಾದ ಉತ್ತರವಿರುತ್ತದೆ.
2. ಪ್ರತಿ ಪ್ರಶ್ನೆಗೆ ಒಂದು ವೃತ್ತವನ್ನು ಮಾತ್ರ ನೀಲಿ ಅಥವಾ ಕಪ್ಪು ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್‌ನಿಂದ ಮಾತ್ರ ತುಂಬತಕ್ಕದ್ದು. ಉತ್ತರವನ್ನು ಮಾರ್ಪಡಿಸಲು ಪ್ರಯತ್ನಿಸಬೇಡಿ.
3. ವೃತ್ತದೊಳಗಿರುವ ಅಕ್ಷರವು ಕಾಣದಿರುವಂತೆ ವೃತ್ತವನ್ನು ಸಂಪೂರ್ಣವಾಗಿ ತುಂಬುವುದು.
4. ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿ ಯಾವುದೇ ಅನಾವಶ್ಯಕ ಗುರುತುಗಳನ್ನು ಮಾಡಬೇಡಿ.
5. ಉತ್ತರಿಸಿದ ಪ್ರಶ್ನೆಗಳ ಒಟ್ಟು ಸಂಖ್ಯೆಯನ್ನು O.M.R. ಹಾಳೆಯಲ್ಲಿ ನಿಗದಿಪಡಿಸಿರುವ ಜಾಗದಲ್ಲಿ ನಮೂದಿಸತಕ್ಕದ್ದು, ಇಲ್ಲವಾದಲ್ಲಿ O.M.R. ಹಾಳೆಯನ್ನು ಮೌಲ್ಯಮಾಪನಕ್ಕೆ ಪರಿಗಣಿಸುವುದಿಲ್ಲ.

Note : English version of the instructions is printed on the front cover of this booklet.